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Modern Chemistry 65 Chapter Test
Chapter: Chemical Equations and
Reactions In the space provided, write
the letter of the term or phrase that best

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completes each statement or best answers each question. _____ 1. You mix solution A with solution B in a beaker. Which of the following observations does not help you prove that a chemical reaction has occurred? a.

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If two or more compounds are composed of elements A and B, the ratio of the masses of B combined with 1 g of A is always a ratio of small whole numbers. This is a statement of the law of Multiple proportions

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Chapter 11 Chemistry Test. reactions in which O_2 is consumed by combining with another substance; always releases heat and/or other forms of energy; produces one or more oxygen-containing compounds; produces CO_2 and H_2O .

Chemical Reaction; Grade 8 Chapter

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Chemical Reactions - Prentice Hall
Chemistry Chapter 11. A representation
of a chemical reaction A chemical
equation that does not indicate the
relative amount... A substance that
speeds up the reaction, but is not used
up in... Small whole numbers that are
placed in front of the formulas i...

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Chemical Equation A representation...

Chapter 7: Chemical Reactions - Practice Test Questions ...

A chemical reaction in which energy is absorbed is called a. synthetic. b.

exothermic. c. endothermic. ____ 4.

Chemical equations must be balanced because chemical reactions obey the

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principle of a. activation. b. conservation of matter. c. decomposition. ____ 5. In a chemical equation, the number in front of a chemical

Assessment Chapter Test A

5. Name four ways you can increase the rate of a chemical reactions. a. Increase temperature c. Increase surface area b.

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Increase concentration d. Add a catalyst
6. Name four clues or indicators that a chemical reaction occurred. a. Explosion
c. Gas produced b. Temperature change
d. Precipitate formed 7. List the five types of chemical reaction: a.

**SCIENCE 8CP Name: Period:
Chemical Reactions Test Review ...**

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The reaction you want to take place will combine the soccer ball and the goal to score a point. The combination of the goal and the ball into something called a point is the product. However, the ball and the goal start meters away from each other. By themselves, they will not combine to score a point.

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Chapter 2 The Chemistry of Life Chapter
Test B. Multiple Choice. Write the letter
that best answers the question or
completes the statement on the line
provided. ____ 1. The three particles that
make up an atom are a. protons,
neutrons, and isotopes. b. neutrons,
isotopes, and electrons. c. positives,

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negatives, and electrons.

sxPS G8 CT ch06-B 067-070

Chapter Tests Chemical Reactions
(continued) 32. A) Enzymes are
catalysts. They speed up chemical
reactions by lowering the activation
energy of reactions needed for life.
Enzymes allow reactions that are

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Answers

necessary for life to occur at body temperature. B) Sawdust has a greater surface area than wood. Increased surface area of a reactant will ...

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Chemical Reactions Practice Test

Chapter 2 Name _____ Date _____ Hour _____

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 Multiple Choice Identify the choice that best completes the statement or answers the question. ____ 1. The only sure evidence for a chemical reaction is
a. the formation of a gas. b. changes in properties. c. the production of one or more new substances.

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reactions Flashcards ...

Online Test of Chapter 1 Chemical Reactions and Equations 1 Science| Class 10th. Q1. A chemical reaction has taken place in which of the following process? (i) Ice melts into water (ii) A wet shirt got dried in sunlight (iii) A brown layer is formed over iron rod kept in air (iv) Sugar getting dissolved in

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water. Q2.

Chapter 2 Chemical Reactions Practice Test - MAFIADOC.COM

5. Consider the reaction: $\text{Mg} + \text{CuCl}_2 = \text{MgCl}_2 + \text{Cu}$ Which type of reaction is shown above? A whole number that appears in front of a compound or element in a balanced chemical

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equation. A whole number that appears as a subscript in front of a formula in a balanced chemical equation. Question 23 23.

Multiple Choice - Lawndale High School

A synthesis reaction is a reaction between at least two compounds in

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Answers

which. a. one breaks down into at least two products. b. a compound is decomposed by an electric current. c. a compound burns in the presence of oxygen. d. a new, more complex compound is formed.

Chapter Test B Chemical Reactions

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Answers

Chapter Test B, continued _____ 7. In a reaction, the ions of two compounds exchange places in aqueous solution to form two new compounds. This reaction is called a

- synthesis reaction.
- decomposition reaction.
- single-displacement reaction.
- double-displacement reaction.

_____ 8. The use of a double arrow in a chemical equation

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Answers

indicates that the reaction a. is reversible.

Chapter 1 Chemical Reactions and Equations MCQ Test 1 ...

In a chemical reaction, the substances that undergo change. The new substances formed as a result of change in a chemical... A representation of a

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Answers

chemical reaction in which the reactants... The numbers that appear in the formula. Reactants In a chemical reaction, the substances that undergo change.

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CHAPTER 8 REVIEW Class SĖc7ĭB-n

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Answers

Chemical Equations and Reactions

SHORT ANSWER Answer the following questions in the space provided. 1.

Match the symbol on the left with its appropriate description on the right. (aq)

(a) (b) (c) (d) (f) A precipitate forms. A gas forms. A reversible reaction occurs. Heat is applied to the reactants.

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Test , Class 10, Chemistry, CBSE-
Chemical Reactions and Equations.

Answers Chemical Reactions magnesium and oxygen in MgO. Mg

35) One sign that a chemical reaction has occurred is the formation of new substances. a) True b) False _____ a. 36)

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Answers

A chemical equation shows a chemical reaction using symbols instead of words.

a) True b) False _____ b. 37) The reactants are the new materials produced during a chemical reaction. a) True b) False

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Answers

A chemical reaction in which an element replaces one element in a compound
double replacement reactions the negative ions in two compounds switch places, forming two new compounds.